AQA Chemistry GCSE Student checklist

C2

Name ..... Class ..... Date .....

The periodic table

Lesson	Target 4		Target 6		Target 8	
C2.1 Development of the periodic table	I can list the significant models for ordering the elements.	☐	I can describe how the elements are arranged in groups and periods in the periodic table.	☐	I can explain how and why the ordering of the elements has changed over time.	☐
	I can state how the elements are ordered in the periodic table.	☐	I can explain why the periodic table was a breakthrough in how to order elements.	☐		
C2.2 Electronic structures and the periodic table	I can define a group and period in the periodic table.	☐	I can describe how the electronic structure of metals and non-metals are different.	☐	I can explain how the electronic structure of metals and non-metals affects their reactivity.	☐
	I can describe how electronic structure is linked to the periodic table.	☐	I can explain in terms of electronic structure how the elements are arranged in the periodic table.	☐	I can use the periodic table to make predictions about the electronic structure and reactions of elements.	☐
	I can state that noble gases are unreactive.	☐	I can explain why the noble gases are unreactive and the trend in their boiling points.	☐	I can predict the electronic structure of stable ions for the first 20 elements.	☐
C2.3 Group 1- the alkali metals	I can name the first three elements in Group 1.	☐	I can recognise trends in supplied data.	☐	I can illustrate the reactions of Group 1 metals with balanced symbol equations.	☐
	I can describe the Group 1 metals as having low densities.	☐	I can explain why the elements in Group 1 react similarly and why the first three elements float on water.	☐	I can explain how Group 1 metals form ions with a +1 charge when they react with non-metals.	☐
	I can write word equations from descriptions of how Group 1 metals react with water.	☐	I can Describe how you can show that hydrogen and metal hydroxides are made when Group 1 metals react with water.	☐	I can justify how Group 1 metals are stored and the safety precautions used when dealing with them.	☐

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C2.4 Group 7- the halogens	I can name the first four elements in Group 7.	☐	I can recognise trends in supplied data.	☐	I can illustrate the reactions of Group 7 metals with balanced symbol equations.	☐
	I can recognise a halogen displacement reaction.	☐	I can explain why the elements in Group 7 react similarly.	☐	I can explain how Group 7 non-metals form ions with a −1 charge when they react with metals.	☐
	I can describe the main properties of halogens.	☐	I can explain how to complete a halogen displacement reaction and explain what happens in the reaction.	☐	I can explain in detail how to compare the reactivity of the Group elements.	☐
C2.5 Explaining trends	I can state the trend in reactivity in Group 1.	☐	I can explain how electronic structure affects the trend in reactivity of Group 1 and Group 7 elements.	☐	I can use electronic structure to explain the trends in physical and chemical properties of Group 1 and Group 7 elements.	☐
	I can state the trend in reactivity in Group 7.	☐	I can use the nuclear model to explain how the outer electrons experience different levels of attraction to the nucleus.	☐	I can apply knowledge of reactivity of Groups 1 and 7 to suggest and explain the trend in reactivity of Group 2 and 6.	☐